pH practice

$pH = -\log[H+]$
$pOH = -\log[OH^-]$
$pH + pOH = 14$
$[H^+]x[OH^-] = 1.0x10^{-14}$

For each of the following problems, identify the pH AND tell if the solution is acidic, basic, or neutral.

1. $[OH^-] = 7.6x10^{-10}$
2. $[H^+] = 1.3x10^{-7}$
3. $pOH = 7.9$
4. $pOH = 2.5$
5. $[OH^-] = 8.8x10^{-2}$
6. $[H^+] = 3.5x10^{-5}$
7. $[OH^-] = 6.6x10^{-3}$
8. $pOH = 7$
9. $[H^+] = 4.0x10^{-1}$
10. $lH^+] = 9.1x10^{-1}$

	$[H^+]$	pH	$[OH^-]$	pOH	Acidic, Basic, or Neutral
11	10^{-5} M	5	10^{-9} M	9	Acidic
12		7			
13			10^{-4} M		
14	10^{-2} M				
15				11	
16		12			
17			10^{-5} M		
18	10^{-11} M				
19				13	
29		6			

Selected Answers:
1. pH = 4.88 ($[H^+]= 1.3x10^{-5}$ M)
3. pH = 6.1
6. pH = 4.46
13. $[H^+] = 10^{-10}$ M; pH = 10; pOH = 4; Basic